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Chapter 11: Chemical Reactions Study Guide. A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two. Small whole number ratios that are placed in front of the formulas in an equation in order to balance it.

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### CHAPTER 11 CHEMICAL REACTIONS GUIDED READING ANSWER KEY ...

Study Guide: Chapter 11—Chemical Reactions Key Concepts  
11.1—Describing Chemical Reactions \*How do you write a word equation? To write a word equation, write the names of the reactants to the left of the arrow separated by a plus sign; write the names of the products to the right of the arrow, also separated by plus signs.

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Access Free Chapter 11 Chemical Reactions Guided Reading Answers Chapter 11 Chemical Reactions Guided 114 Guided Reading and Study Workbook CHAPTER 11, Chemical Reactions(continued) 9. Use the symbols in Table 11.1 on page 323 to write a skeleton equation for the following chemical reaction. Hydrochloric acid reacts with zinc to produce aqueous

### Chapter 11 Chemical Reactions Guided Reading Answers

Chapter 11 Chemical Reactions 113 SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329) This section explains how to write equations describing chemical reactions using appropriate symbols. It also describes how to write balanced chemical equations when given the names or formulas of the reactants and products in a chemical reaction.

### SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329)

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Section 11.1 - Describing Chemical Reactions In a chemical reaction, the reactants are written on the left and the products on the right. The arrow that separates them is called yield.

Reactants  $\rightarrow$  Products

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical equation An expression that gives the identities and quantities of the ...

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simple chemical reactions. Question 3.

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This chapter discusses the thermodynamics of mixing processes and processes described by reaction equations (chemical equations). It introduces the important concepts of molar mixing and reaction quantities, advancement, and the thermodynamic equilibrium constant.

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Chapter 11 Chemical Reactions Guided Reading Answers Guided Reading and Study Workbook CHAPTER 11, Chemical Reactions(continued) 9 Use the symbols in Table 111 on page 323 to write a skeleton equation for the following chemical reaction Hydrochloric acid reacts with zinc to produce aqueous zinc(II) chloride and hydrogen

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